

Course No: DMED 1207  
Course Title: General Chemistry  
Date: 28 / 11 /2017  
No. of Questions: (6)  
Time: 1 hours  
Using Calculator (Yes)

University of  
Palestine  
  
Second Exam  
2017/2018

Instructor Name:  
Student No.: \_\_\_\_\_  
Student Name: \_\_\_\_\_  
College Name: \_\_\_\_\_  
Dep. / Specialist: \_\_\_\_\_  
Using Dictionary (No)  
Total Grade: 15

**Question I. Define the following terms " 2 Marks"**

**1. Un saturated Solution**

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**2. Reversible reaction**

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**3. Weak Base**

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**4. Oxidation**

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**Question II. True Or False**

**( 2.5 Marks )**

1. ( ) When calculating the oxidation number, the total charge of the compound is equal zero .
2. ( ) Acids reacts with metals to give CO<sub>2</sub> .
3. ( ) Strong acids are a good electricity conductor .
4. ( ) Irreversible reaction occurs in weak acid and base .
5. ( ) In free elements each atom has an oxidation number of zero.

**Question III. Choose a correct answer**

**( 2.5 Marks )**

1. How would you prepare 5% (w/v) glucose solution in 150 ml solution ?  
a) 15g glucose  
b) 15ml glucose  
c) 7.5 g glucose +142.5 g distilled H<sub>2</sub>O  
d) 100ml g distilled H<sub>2</sub>O
2. What is the concentration of a dilute solution his volume is 350 ml starting with a stock solution of 10 M HCL and volume 150 ml  
a) 4.3 L      b) 0.01ml      c) 0.24M      d) 4.3 M
3. reaction of some metals with acids gives

- a)  $H_2O$       b)  $H_2$       c)  $O_2$       d)  $N_2$

4. The element which takes electrons is

- a) reducing agent      b) oxidizing agent      c) buffer      d) solvent

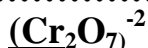
5. What is the molarity 1.5 g KOH, was dissolved in 100 mL water? F.wt KOH=56

- a) 0.27 M                      b) 0.27 mol / L  
 c) 0.27 m                      c) a and b

**Question IV. Find the oxidation number of the under lined element ( 2 Marks )**



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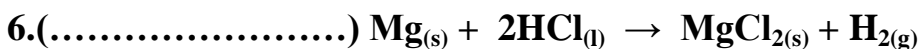
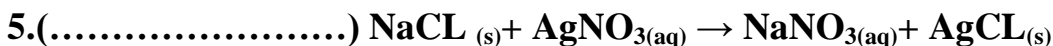
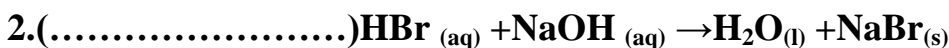
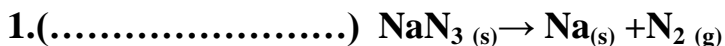


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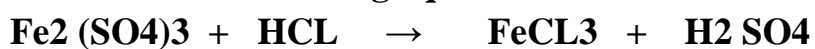
**Question V . \_What is the type of these reactions ( 3 Marks )**



Question VI . \_ Solve the following

( 3 Marks )

a. Balance the following equation and find



b . How many grams of FeCL3 will be produced from the complete reaction of 1mole of Fe2 (SO4)3

c. how many grams of sulfur will be produced from the last equation?

atomic w.t of Fe = 56g

atomic w.t of CL = 35.5g

atomic w.t of S = 32g

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**End of Questions**  
**Good Luck**