

Course No:
Course Title: Gen.Chem
Date:27/11/2017
No. of Questions: (1)
Time: 1 hour
Using Calculator (yes)

University of Palestine



First Exam.
2017-2018
Total Grade:15

Instructor Name: Dr. Asmaa
Student No.: _____
Student Name: _____
College Name: _____
Dep. / Specialist:
Using Dictionary (No)

Question One:

1. Which of the following is a representation of the set conditions known as standard temperature and pressure (STP)?

- a) 0 K and 1 atm
- b) 0°F and 760 torr
- c) 0°C and 760 atm
- d) 273°C and 1 atm
- e) 273 K and 760 torr

2. Which of the following gases has the highest average velocity at a given temperature

- a) Oxygen
- b) Carbon monoxide
- c) Neon
- d) Sulfur dioxide
- e) Hydrogen chloride

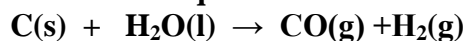
3. A gas has a density of 2.68 g/L (STP). What is the gas?

- a) CO₂
- b) SO₂
- c) NO₂
- d) COS
- e) He

4. A 2.0 L quantity of 0.10 M HCl contains

- a) 1.0 mol of HCl
- b) 0.20 mol of H₂O
- c) 20 mol of HCl
- d) 0.05 mol of H₂O
- e) 0.20 mol of HCl

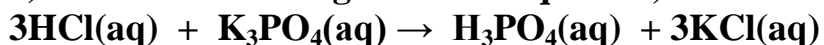
5. Given the equation



what volume of gas measured at STP would be produced from 24.0 g of carbon?

- a) 22.4 L
- b) 89.6 L
- c) 44.8 L
- d) 11.2 L
- e) 4.0 L

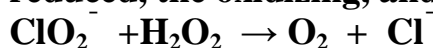
1) Given the following balanced equation,



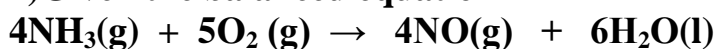
What mass of K_3PO_4 is needed to react with 325 mL of 0.250 M HCl?

2) Write the balanced molecular, total ionic, and net ionic equations illustrating the neutralization of HNO_3 with $\text{Sr}(\text{OH})_2$.

3) In the following reaction, indicate the substance oxidized, the substance reduced, the oxidizing, and the reducing agent

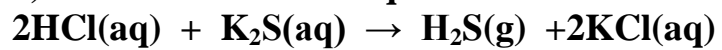


4) Given the balanced equation



What volume of NO gas measured at 550 torr and 25°C will be produced from 19.5g O_2 ?

5) Given the balanced equation



What volume of H_2S measured at STP would be evolved from 1.56 L of 0.552 M HCl solution with excess K_2S present?