

Course No: DMED 1207  
Course Title: General Chemistry  
Date: 29 / 11 /2016  
No. of Questions: (6)  
Time: 1 hours  
Using Calculator (Yes)

University of  
Palestine  
  
Second Exam  
2016/2017

Instructor Name:  
Student No.: \_\_\_\_\_  
Student Name: \_\_\_\_\_  
College Name: \_\_\_\_\_  
Dep. / Specialist: \_\_\_\_\_  
Using Dictionary (No)  
Total Grade: 15

**Question I. Define the following terms " 3 Marks"**

**1. Saturated Solution**

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**2. Irreversible Reaction**

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**3. Weak Base**

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**4. Reducing Agent**

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**5. Oxidation**

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**6. Dilution**

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**Question II. True Or False**

**( 3 Marks )**

1. ( ) Acids reacts with  $\text{Na}_2\text{CO}_3$  to give  $\text{H}_2$  (g).
2. ( ) The Concentration depends on the quantity of solute if the volume of solution is fixed
3. ( ) Strong acids are a good electricity conductor .
4. ( ) Ionization of a strong acids and bases in water is weak.
5. ( ) water is bad solvent for ionic compound .
6. ( ) buffer solution consist of a weak base and its conjugate acid.

**Question III. Choose a correct answer**

**( 3 Marks )**

1. The PH of water should be around  
a . 2                    b . 5                    c . 7                    d . 12
2. What is the concentration of a dilute solution his volume is 500 ml starting with a stock solution of 10 M Hcl and volume 3 ml  
a . 30 L                b . 0.06ml             c . 150M                d . 0.06 M
3. reaction of some metals with acids gives  
a. H<sub>2</sub>O                b . H<sub>2</sub>                c . O<sub>2</sub>                d . N<sub>2</sub>
4. The element which takes electrons is  
a . reducing agent    b . oxidizing agent    c . buffer            d . solvent
5. Reversible reaction occurs in :  
a . strong acids                b . strong base                c . weak acid and bases  
d . none of the above
6. 5% Nacl aqueous solution has volume of  
a. 5ml                b.100ml                c. 5g                d. 50ml

**Question IV. Find the oxidation number of the under lined element**

**( 2 Marks )**



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**Question V. Do as requested ( 1 Marks )**

**1- How many grams of sodium carbonate (Na<sub>2</sub>CO<sub>3</sub>) are required to prepare 350 ml solution with a concentration of 1.5 M ?**

**a.wt Na= 23 g , a.wt C= 12g , a.wt O= 16g**

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**A. Question VI . \_What is the type of these reactions ( 3 Marks )**

- 1.(.....)  $\text{MgCl}_2\text{(s)} + \text{FeSO}_4\text{(aq)} \rightarrow \text{MgSO}_4\text{(aq)} + \text{FeCl}_2\text{(s)}$
- 2.(.....)  $\text{N}_2\text{O}_5\text{(g)} + \text{H}_2\text{O(l)} \rightarrow 2\text{HNO}_3\text{(aq)}$
- 3.(.....)  $\text{NH}_4\text{Cl} \rightarrow \text{NH}_3 + \text{HCl}$
- 4.(.....)  $\text{KOH(aq)} + \text{CH}_3\text{COOH(aq)} \rightarrow \text{CH}_3\text{COOK(s)} + \text{H}_2\text{O}$
- 5.(.....)  $\text{CuSO}_4\text{(aq)} + 2\text{NaOH(aq)} \rightarrow \text{Cu(OH)}_2\text{(s)} + \text{Na}_2\text{SO}_4\text{(aq)}$
- 6.(.....)  $\text{Mg(s)} + 2\text{HCl(l)} \rightarrow \text{MgCl}_2\text{(s)} + \text{H}_2\text{(g)}$

***End of Questions  
Good Luck***