

Course No: DNTS1302
Course Title: General Chemistry
Date: 27/11/2014
No. of Questions: (6)
Time: 1 hours
Using Calculator (No)

University of Palestine



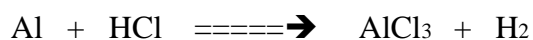
Mid term Exam
2014-2015
Total Grade:

Instructor Name: _____
Student No.: _____
Student Name: _____
College Name: _____
Dep. / Specialist: _____
Using Dictionary (No)

Question One: (4 marks)

**You may use calculators in
this exam. Molar gas
volume: 22.4 L/mol.
Avogadro Nr: 6,022x10(23)**

Calculate the **density** of hydrogen gas produced as 30ml of 0.2M HCl reacts with **excess** Aluminum.
Make sure that the equation is *balanced* first.



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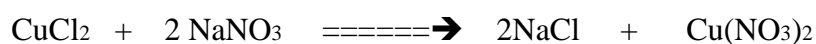


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Question Two: (4 marks)

When 15g of CuCl_2 reacted with 20g of sodium nitrates; 10g of sodium chloride was produced, calculate the *percentage yield* of sodium chloride:



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Question Three: (3 marks)

A compound is made of 43.6% phosphorous, 56.4% oxygen. If its relative molecular mass is 284 find its *molecular formula*.

Question Four: (2 marks)

4) State all *practical* steps involved to prepare 50 mL of 0.4M **ammonium nitrate**.

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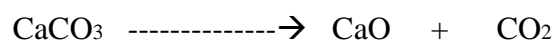
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Question Five: (5 marks)

A rock weighing 95.0 g, when heated produced 8.6 litres of carbon dioxide.

Calculate the **percentage** of calcium carbonate contained in the rock and the **density** of CO₂:



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Question Six: (2 marks)

- 1) Find the **number** of hydrogen atoms in 20.0 g of glucose, $C_6H_{12}O_6$.

End of Questions
Good Luck